

Page 775. Table III. Heading of last column should be $-\Delta S^\ddagger$.

D. J. Pasto and K. Garves: The Silver Ion Assisted Solvolysis of Phenacyl Halides in Aqueous Ethanol.

Page 778. Table III. Heading of last column should be $-\Delta S^\ddagger$.

Robert E. Buckles, James L. Miller, and Roland J. Thurmaier: The Kinetics of the Addition of Bromine to *cis*- and *trans*-Stilbene.

Page 891. Column 1. the equilibrium of eq 14 may very well have a part in the reaction mechanism of the addition of bromine to *cis*- or *trans*-stilbene in the presence of the tetrabutylammonium tribromide, but its involvement was incorrectly interpreted in the text of the article. If the ion aggregate formed had geometry unfavorable to the formation of dibromide product the equilibrium would lead to a slower reaction, but it would not affect the rate law derived from the mechanism. On the other hand, were this equilibrium an intermediate step of the product-forming reaction of eq 13 it would not appear in the rate law (eq 15) unless all of the reaction of eq 13 were treated in a step-wise fashion. In any event eq 15 should be corrected to read

$$-d[\text{Br}_2]/dt = \frac{k_3 k_4 k_{13} [\text{E}] [\text{Br}_2]^2 [\text{R}_4\text{NBr}_3]}{k_{-3} k_{-4} + k_{-3} k_{13} [\text{R}_4\text{NBr}_3]}$$

If the formation of a higher ion aggregate involving two molecules of tribromide salt were tying up the intermediate ion in an unreactive form, it also would *not* affect the rate law as described in the text. If this reaction were an intermediate step in the formation of product the rate law would have the form of the corrected eq 15 except that the $[\text{R}_4\text{NBr}_3]$ term would be squared in both numerator and denominator. Either of these corrected rate laws would explain the leveling effect at higher bromide concentrations but *not* the maximum value of the rate constant. Such a maximum value of the rate constant could be explained by reversible inhibition involving two or more molecules of tribromide salt but not involving an intermediate of the product-forming reaction sequence.

Johannes Meienhofer: Synthesis of Peptide Derivatives with Actinomycin D Sequence.

Page 1143. Formula IX (actinomycin D acid). The phenoxazone ring system carries a nitro group in position 2. Correctly it should be an amino group.

Howard Peters, Robert A. Archer, and Harry S. Mosher: 2,3-Dimethyl-1,2,3,4-tetrahydronaphthalenes and 4,5-Dimethylcyclohexenes. Proton Magnetic Resonance and Stereochemistry.

Page 1383. Column 2, line 6. "*cis*-3,4-dimethylcyclohexene" should read "*cis*-4,5-dimethylcyclohexene."

Page 1383. Column 2, line 18. "... in going from the *cis* to *trans* from ..." should read "... in going from the *trans* to *cis* form ..."

Page 1384. Column 2, lines 5 and 6. "2,3-dimethyl-2,3,4-tetrahydronaphthalenes" should read "2,3-dimethyl-1,2,3,4-tetrahydronaphthalenes."

Page 1387. Column 1, line 20. "*cis*-3,4-Dimethylcyclohexene" should read "*cis*-4,5-Dimethylcyclohexene."

Page 1387. Column 1, ref 40. Page 29 should read 93.

Page 1387. Column 2, line 5. "*trans*-3,4-Dimethylcyclohexene" should read "*trans*-4,5-Dimethylcyclohexene."

Ernest Lustig and Edward P. Ragelis: Proton Spin Coupling in 1-Indanone.

Pages 1399-1400. The figure appearing with the caption for Figure 2 should be Figure 4; the figure appearing with the caption for Figure 3 should be Figure 2; and the figure appearing with the caption for Figure 4 should be Figure 3.

K. Darrell Berlin and Robert B. Hanson: Ring Cleavage of 2,2,4-Trimethyl-3-hydroxy-3-pentenoic Acid β -Lactone by the Anion of Diisopropyl Ketone.

Page 1764. Scheme I. Under structure 3 is an arrow with a small 6. The 6 is incorrect, and should be 1.

Gerald A. Selter and Kirk D. McMichael: Displacement and Elimination Reactions of 5 α ,6 α -Epoxy-3 β -cholestanyl *p*-Toluenesulfonate in Dimethylformamide.

Page 2547. Structure 3a, X = COOH should read X = HCOO.

Leo A. Paquette and William C. Farley: The Chloramine-Induced Oxidative Dimerization of Phenols.

Page 2723. Reference 35 should read "G. R. Yohe, J. E. Dunbar, R. L. Pedrotti, F. M. Scheidt, F. G. H. Lee, and E. C. Smith. . . ."

Lars Skattebøl, Bernice Boulette, and Stanley Solomon: Reactions of sulfides with *t*-Butyl Hypochlorite.

Page 3112. Scheme I should appear as shown below.

